

# **Chemistry A**

Advanced Subsidiary GCE

Unit **F321**: Atoms, Bonds and Groups

## **Mark Scheme for June 2012**

---

OCR (Oxford Cambridge and RSA) is a leading UK awarding body, providing a wide range of qualifications to meet the needs of candidates of all ages and abilities. OCR qualifications include AS/A Levels, Diplomas, GCSEs, OCR Nationals, Functional Skills, Key Skills, Entry Level qualifications, NVQs and vocational qualifications in areas such as IT, business, languages, teaching/training, administration and secretarial skills.

It is also responsible for developing new specifications to meet national requirements and the needs of students and teachers. OCR is a not-for-profit organisation; any surplus made is invested back into the establishment to help towards the development of qualifications and support, which keep pace with the changing needs of today's society.

This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which marks were awarded by examiners. It does not indicate the details of the discussions which took place at an examiners' meeting before marking commenced.

All examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes should be read in conjunction with the published question papers and the report on the examination.

OCR will not enter into any discussion or correspondence in connection with this mark scheme.













© OCR 2012

Any enquiries about publications should be addressed to:

OCR Publications  
PO Box 5050  
Annesley  
NOTTINGHAM  
NG15 0DL

Telephone: 0870 770 6622  
Facsimile: 01223 552610  
E-mail: [publications@ocr.org.uk](mailto:publications@ocr.org.uk)

## Annotations

Annotation	Meaning
	Benefit of doubt given
	Contradiction
	Incorrect response
	Error carried forward
	Ignore
	Not answered question
	Benefit of doubt not given
	Power of 10 error
	Omission mark
	Rounding error
	Error in number of significant figures
	Correct response

**Subject-specific Marking Instructions**

Abbreviations, annotations and conventions used in the detailed Mark Scheme (to include abbreviations and subject-specific conventions).

<b>Annotation</b>	<b>Meaning</b>
<b>DO NOT ALLOW</b>	Answers which are not worthy of credit
<b>IGNORE</b>	Statements which are irrelevant
<b>ALLOW</b>	Answers that can be accepted
( )	Words which are not essential to gain credit
<u>    </u>	Underlined words must be present in answer to score a mark
<b>ECF</b>	Error carried forward
<b>AW</b>	Alternative wording
<b>ORA</b>	Or reverse argument

The following questions should be annotated with ticks to show where marks have been awarded in the body of the text: **1(e)(i), 2(b), 3(b)(ii)**

Question		Answer	Marks	Guidance
1	(a)	<p>The (weighted) mean <b>mass</b> of an <b>atom</b> (of an element)  <b>OR</b>  The (weighted) average <b>mass</b> of an <b>atom</b> (of an element)  ✓</p> <p>compared with 1/12th (the mass) ✓</p> <p>of (one atom of) carbon-12 ✓</p>	3	<p><b>ALLOW</b> average atomic mass  <b>DO NOT ALLOW</b> mean mass of an element  <b>ALLOW</b> mean mass of isotopes <b>OR</b> average mass of isotopes  <b>DO NOT ALLOW</b> the singular; 'isotope'</p> <p>For second <b>and</b> third marking points  <b>ALLOW</b> compared with (the mass of) carbon-12 which is 12</p> <p><b>ALLOW</b> mass of <b>one mole</b> of <b>atoms</b> ✓  compared to 1/12th ✓  (mass of) <b>one mole OR 12g</b> of carbon-12 ✓</p> <p><b>ALLOW</b>  <u>          mass of <b>one mole</b> of atoms          </u>  1/12th mass of <b>one mole OR 12g</b> of carbon-12</p>
	(b)	<p><b>FIRST CHECK THE ANSWER ON THE ANSWER LINE</b>  <b>If answer = 32.09 award 2 marks</b></p> $\frac{32 \times 95.02 + 33 \times 0.76 + 34 \times 4.22}{100}$ <p><b>OR</b></p> $30.4064 + 0.2508 + 1.4348$ <p><b>OR</b></p> $= 32.092 \text{ (calculator value) } \checkmark$ <p>(A, =) 32.09 ✓</p>	2	<p><b>ALLOW</b> one mark for ECF from transcription error in first sum provided final answer is to 2 decimal places and is between 32 and 34 and is a correct calculation of the transcription</p> <p>Answer must be 2 decimal places</p>

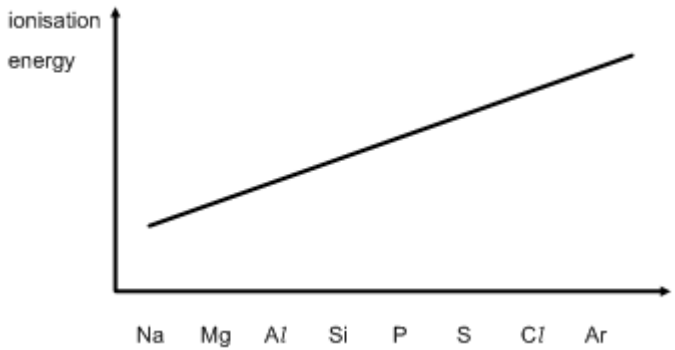
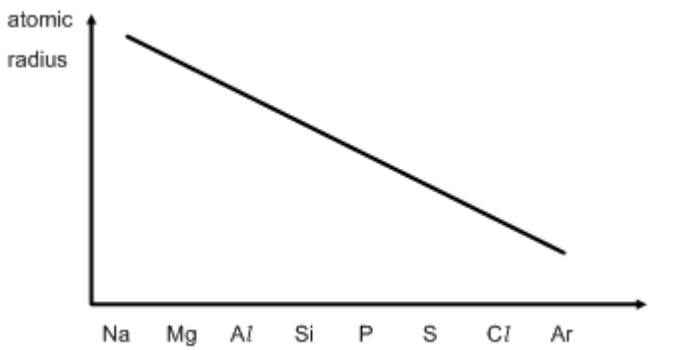
Question		Answer	Marks	Guidance															
1	(c)	<table border="1"> <thead> <tr> <th></th> <th>protons</th> <th>neutrons</th> <th>electrons</th> <th></th> </tr> </thead> <tbody> <tr> <td><math>^{33}\text{S}</math></td> <td>16</td> <td>17</td> <td>16</td> <td>✓</td> </tr> <tr> <td><math>^{34}\text{S}^{2-}</math></td> <td>16</td> <td>18</td> <td>18</td> <td>✓</td> </tr> </tbody> </table>		protons	neutrons	electrons		$^{33}\text{S}$	16	17	16	✓	$^{34}\text{S}^{2-}$	16	18	18	✓	2	Mark by row
	protons	neutrons	electrons																
$^{33}\text{S}$	16	17	16	✓															
$^{34}\text{S}^{2-}$	16	18	18	✓															
	(d)	<p><b>FIRST CHECK THE ANSWER ON THE ANSWER LINE</b>  <b>If answer = <math>5.78 \times 10^{22}</math> award 2 marks</b></p> <p>(mol of atoms) = <math>0.0120 \times 8 = 0.0960</math> (mol)  <b>OR</b>          (no. of molecules) = <math>0.0120 \times 6.02 \times 10^{23} = 7.224 \times 10^{21}</math>  <b>OR</b>          (no. of S atoms in 1 mole of <math>\text{S}_8</math>) = <math>8 \times 6.02 \times 10^{23} = 4.816 \times 10^{24}</math> ✓</p> <p>Correctly calculates (number of atoms) = <math>0.0120 \times 8 \times 6.02 \times 10^{23}</math>  <math>= 5.78 \times 10^{22}</math> (atoms) ✓</p>	2	<p>If there is an alternative answer, check to see if there is any ECF credit possible using working below</p> <p><b>ALLOW</b> <math>5.8 \times 10^{22}</math> up to calculator value of <math>5.7792 \times 10^{22}</math>  <b>ALLOW</b> correct rounding of ECF to 2 significant figures or more up to calculator value  <b>ALLOW</b> answers in non standard form such as <math>0.578 \times 10^{23}</math> correctly rounded to 2 or more significant figures</p>															
	(e) (i)	<p><b>Creating the dipole mark</b>          Uneven distribution of electrons ✓</p> <p><b>Type of dipole mark</b>          Creates or causes an instantaneous dipole <b>OR</b> temporary dipole (in a molecule) ✓</p> <p><b>Induction of a second dipole mark</b>          Causes induced dipoles in neighbouring molecules ✓</p>	3	<p><b>Use annotations with ticks, crosses, ECF etc for this part</b>  <b>ALLOW</b> movement of electrons  <b>ALLOW</b> changing electron density</p> <p><b>ALLOW</b> 'transient', 'oscillating' 'momentary' 'changing'  <b>DO NOT ALLOW</b> induces a temporary dipole for the second marking point</p> <p><b>ALLOW</b> induces a dipole in neighbouring molecules  <b>ALLOW</b> causes a resultant dipole in other molecules  <b>ALLOW</b> atoms for molecules</p>															

Question			Answer	Marks	Guidance
1	(e)	(ii)	Only one type of atom <b>OR</b> No (permanent) dipoles <b>OR</b> non-polar <b>OR</b> no polar bonds ✓	1	<b>ALLOW</b> no difference in electronegativity <b>IGNORE</b> 'No hydrogen bonding' <b>IGNORE</b> 'No lone pairs'
	(f)		+ 2 ✓	1	<b>ALLOW</b> 2(+)
	(g)	(i)	There are no waters of crystallisation ✓	1	<b>ALLOW</b> 'without water' 'no water' etc <b>IGNORE</b> dehydrated
		(ii)	248.2 ✓	1	<b>IGNORE</b> units <b>DO NOT ALLOW</b> 248
		(iii)	<b>FIRST CHECK THE ANSWER ON THE ANSWER LINE</b> <b>If answer = 7.91 (g) award 2 marks</b>  (amount of $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$ ) = $12.41/248.2$ <b>OR</b> = $0.05(00)$ (mol) ✓  (mass of $\text{Na}_2\text{S}_2\text{O}_3$ ) = $0.05 \times 158.2 = 7.91$ (g) ✓	2	If there is an alternative answer, check to see if there is any ECF credit possible using working below  <b>ALLOW</b> ECFs from answer to (g)(ii) for both marking points  <b>ALLOW</b> ECF for calculated mol of $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O} \times 158.2$ correctly calculated for the 2nd mark  <b>ALLOW</b> calculator value or rounding to 3 significant figures or more but <b>IGNORE</b> 'trailing' zeroes, eg 0.200 allowed as 0.2

Question			Answer	Marks	Guidance
1	(h)	(i)	Sulfur has six bonded pairs (and no lone pairs) ✓  Electron pairs repel (one another equally) ✓	2	<b>ALLOW</b> 'It has six bonded pairs' <b>ALLOW</b> bonds for bonded pairs <b>IGNORE</b> regions <b>OR</b> areas of negative charge  <b>ALLOW</b> 'bonds repel' <b>DO NOT ALLOW</b> 'Atoms repel' or 'electrons repel'  'Lone pairs repel more than bonded pairs' would score the second mark but would contradict the first mark if there is no reference to no lone pairs
		(ii)	The ability of an atom to attract electrons ✓ in a (covalent) bond ✓  (The octahedral shape) is symmetrical ✓	3	<b>ALLOW</b> dipoles cancel out <b>IGNORE</b> polar bonds repel <b>IGNORE</b> charges cancel
			<b>Total</b>	<b>23</b>	



Question		Answer	Marks	Guidance
2	(a)	Periodicity ✓	1	<b>ALLOW</b> phonetic versions
	(b)	<p><b>Al bonding mark</b> Al has metallic (bonding) <b>OR</b> has (electrostatic) attraction between positive ions and (delocalised) electrons ✓</p> <p><b>Si bonding mark</b> Si has covalent (bonding) <b>OR</b> has shared pairs of electrons between atoms ✓</p> <p><b>P bonding mark</b> P has <b>induced</b> dipoles <b>OR</b> has van der Waals' forces (between molecules) ✓</p> <p><b>Structure mark 1</b> Al <b>AND</b> Si are Giant ✓</p> <p><b>Structure mark 2</b> P is Simple molecular <b>OR</b> simple covalent ✓</p> <p><b>Bond strength mark</b> Metallic <b>AND</b> covalent are stronger than vdWs <b>OR</b> Bonds broken in Al <b>AND</b> in Si are stronger than the forces broken in P <b>OR</b> More energy is needed to overcome bonds in Al <b>AND</b> Si than the forces in P ✓</p>	6	<p><b>Use annotations with ticks, crosses, ECF etc for this part</b></p> <p><b>DO NOT ALLOW</b> marking point 1 if Al has dipoles <b>OR</b> intermolecular forces <b>OR</b> molecules <b>OR</b> atoms <b>OR</b> attraction between nuclei and electrons <b>OR</b> attraction between oppositely charged ions</p> <p><b>DO NOT ALLOW</b> marking point 2 if Si has dipoles <b>OR</b> intermolecular forces <b>OR</b> molecules but <b>IGNORE</b> 'molecule'</p> <p>Must be induced dipoles <b>ALLOW</b> vdW for van der Waals' <b>IGNORE</b> P has covalent bonds for marking point 3</p> <p><b>Quality of Written Communication: 'giant' spelled correctly once and used in context for the fourth marking point</b></p> <p><b>DO NOT ALLOW</b> covalent bonds are broken in phosphorus for marking point 6, but <b>ALLOW</b> answers that inform Al and Si are stronger than P, ignoring incorrect forces or bonds used above <b>IGNORE</b> 'heat' but <b>ALLOW</b> 'heat energy'</p>

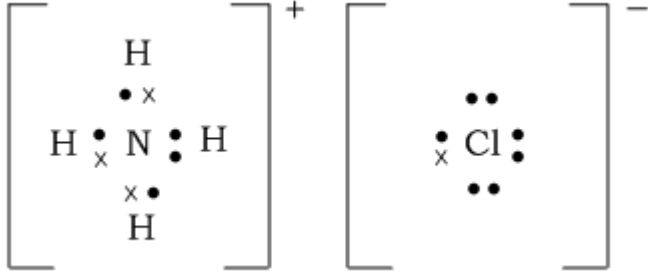
Question			Answer	Marks	Guidance
2	(c)	(i)	Increasing straight line <b>OR</b> curve from Na to Ar ✓ 	1	<b>ALLOW</b> bar charts <b>OR</b> points <b>IGNORE</b> the standard of drawing as long as the trend is clear <b>IGNORE</b> decrease between Mg/Al and P/S  Essentially the mark is for Na < Mg < Si < P < Cl < Ar <b>AND</b> Al < Si <b>AND</b> S < Cl
		(ii)	Decreasing straight line <b>OR</b> curve from Na to Ar ✓ 	1	<b>ALLOW</b> bar charts <b>OR</b> points <b>IGNORE</b> the standard of drawing as long as the trend is clear <b>IGNORE</b> Ar  Essentially the mark is for Na > Mg > Al > Si > P > S > Cl
<b>Total</b>				<b>9</b>	

Question		Answer	Marks	Guidance
3	(a)	$(1s^2) 2s^2 2p^6 3s^2$ ✓	1	<b>IGNORE</b> $1s^2$ seen twice <b>ALLOW</b> subscripts
	(b) (i)	$Mg^+(g) \rightarrow Mg^{2+}(g) + e^-$ Equation correct ✓ State symbols correct ✓	2	<b>ALLOW</b> $Mg^+(g) - e^- \rightarrow Mg^{2+}(g)$ for 2 marks The second mark is dependent upon the first mark except for the following close attempts for the first mark: <b>ALLOW</b> the following for one mark as states are correct $Mg(g) \rightarrow Mg^{2+}(g) + 2e^-$ $Mg(g) + e^- \rightarrow Mg^{2+}(g) + 2e^-$ <b>ALLOW</b> e for electron <b>IGNORE</b> states on electron
	(ii)	<b><i>Ionic radius mark</i></b> $Mg^{(+)}$ has smaller (ionic) radius <b>OR</b> has less shells ✓  <b><i>Shielding mark</i></b> (outermost electron) of $Mg^{(+)}$ experience less shielding ✓  <b><i>Nuclear attraction mark</i></b> More nuclear attraction on (outermost electrons) <b>OR</b> Outer electrons are attracted more strongly (to the nucleus) ✓  <b>ORA</b> throughout	3	<b>Use annotations with ticks, crosses, ECF etc for this part</b>  <b>ALLOW</b> $Mg^{(+)}$ has less energy levels <b>ALLOW</b> $Mg^{(+)}$ has electrons in lower energy level <b>ALLOW</b> $Mg^{(+)}$ has electrons closer to nucleus <b>IGNORE</b> $Mg^{(+)}$ has less orbitals <b>OR</b> less sub-shells <b>IGNORE</b> atomic for ionic <b>IGNORE</b> 'different shell'  <b>ALLOW</b> screening for shielding <b>ALLOW</b> $Mg^{(+)}$ has <b>less</b> electron repulsion from inner shells  <b>Quality of Written Communication: 'nuclear' OR 'nucleus' OR 'electron(s)' spelled correctly once and used in context for the third marking point</b>  <b>ALLOW</b> $Mg^{(+)}$ has more nuclear pull <b>IGNORE</b> $Mg^{(+)}$ has more effective nuclear charge <b>DO NOT ALLOW</b> more nuclear charge for more nuclear attraction for the third mark

Question			Answer	Marks	Guidance
3	(c)	(i)	Sr <sup>2+</sup> ✓ OH <sup>-</sup> ✓	2	<b>ALLOW</b> 2OH <sup>-</sup> <b>ALLOW</b> 2 marks for Sr(OH) <sub>2</sub> → Sr <sup>2+</sup> + 2OH <sup>-</sup> <b>ALLOW</b> 1 mark for Sr <sup>2+</sup> + 2OH <sup>-</sup> → Sr(OH) <sub>2</sub> <b>IGNORE</b> H <sup>+</sup>
		(ii)	Sr has lost (two) electrons ✓	1	<b>ALLOW</b> Sr → Sr <sup>2+</sup> + 2e <sup>-</sup> <b>IGNORE</b> references to oxidation numbers
		(iii)	SrO <b>AND</b> H <sub>2</sub> O ✓	1	<b>ALLOW</b> acceptable alternatives from Sr salts and alkalis eg SrCl <sub>2</sub> + NaOH
	(d)	(i)	It shows the oxidation number of the sulfur <b>OR</b> the name without the IV is ambiguous ✓	1	<b>DO NOT ALLOW</b> 'the <b>charge</b> on sulfur' <b>DO NOT ALLOW</b> 'shows the oxidation number of the sulfate' <b>ALLOW</b> Otherwise it could be SrSO <sub>4</sub> <b>ALLOW</b> Sulfur has different oxidation numbers AW
		(ii)	H <sub>2</sub> SO <sub>3</sub> ✓	1	
<b>Total</b>				<b>12</b>	

Question			Answer	Marks	Guidance
4	(a)	(i)	$\text{Cl}_2 + \text{H}_2\text{O} \rightarrow \text{HClO} + \text{HCl} \checkmark$	1	
		(ii)	(Chlorine compounds are) carcinogenic <b>OR</b> (Chlorine compounds are) toxic <b>OR</b> poisonous $\checkmark$	1	<b>ALLOW</b> 'they' <b>OR</b> 'chlorinated hydrocarbons' <b>OR</b> 'it' for 'chlorine compounds'  <b>IGNORE</b> harmful <b>OR</b> dangerous <b>IGNORE</b> references to HCl or HClO <b>IGNORE</b> chlorine is toxic  <b>DO NOT ALLOW</b> chlorine is carcinogenic
	(b)	(i)	Precipitation $\checkmark$	1	
		(ii)	$\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s}) \checkmark$	1	Equation <b>AND</b> state symbols required for mark <b>DO NOT ALLOW</b> spectator ions
	(c)	(i)	$8.604/143.4 = 0.06(00) \text{ (mol)} \checkmark$	1	

Question		Answer	Marks	Guidance
4	(c) (ii)	<p>If a Group 2 chloride is used  amount of Group 2 chloride = <math>\frac{1}{2} \times 0.0600</math> <b>OR</b> = 0.0300 mol ✓</p> <p>Mass of 1 mol of Group 2 chloride  = <u>2.86</u> = 95.3(3) ✓  0.0300</p> <p>[Relative atomic mass of M = 95.3(3) – 71.0) = 24.3 (g mol<sup>-1</sup>)] <b>AND</b>  metal = Mg ✓</p>	3	<p><b>DO NOT ALLOW</b> 24.3 and Mg without appropriate working</p> <p>Check to see if there is any ECF credit possible using working below</p> <p><b>ALLOW</b> calculator value or rounding to 2 significant figures or more but <b>IGNORE</b> 'trailing' zeroes, eg 0.200 allowed as 0.2</p> <p><b>ALLOW</b> ECF for correctly calculated <math>\frac{1}{2} \times</math> answer to (c)(i)</p> <p>Must be at least 1 decimal place for second marking point</p> <p><b>ALLOW</b> ECF for 2.86/mol of metal chloride seen above  eg MCl will give 0.0600 mol of metal chloride and this will likely give 2.86/0.0600 = 47.7  eg <math>MCl_3</math> will give 0.0200 mol of metal chloride and this will likely give 2.86/0.0200 = 143.0</p> <p><b>ALLOW</b> ECF for mass of Group 2 chloride – 71.0 provided it is not a negative value</p> <p><b>ALLOW</b> ECF even if molar mass of chloride was given as a whole number above</p> <p><b>ALLOW</b> ECF for mass of metal chloride – 35.5 if amount of metal chloride = 0.0600 mol  eg 47.7 – 35.5 = 12.2 <b>AND</b> Be</p> <p><b>ALLOW</b> ECF for mass of metal chloride – 106.5 if amount of metal chloride = 0.0200 mol  eg 143.0 – 106.5 = 36.5 <b>AND</b> Ca</p>

Question		Answer	Marks	Guidance
4	(d) (i)	A shared pair of electrons <b>AND</b> both electrons are donated by one atom ✓	1	
	(ii)	$\text{NH}_4^+$ <b>AND</b> $\text{Cl}^-$ ✓	1	<b>ALLOW</b> $\text{NH}_4\text{Cl} \rightarrow \text{NH}_4^+ + \text{Cl}^-$ <b>OR</b> $\text{NH}_4^+ + \text{Cl}^- \rightarrow \text{NH}_4\text{Cl}$
	(iii)	Ammonium ion with three covalent ' <i>dot-and-cross</i> ' bonds <b>AND</b> one dative covalent bond ✓  Chloride ion with $8e^-$ <b>AND</b> 1 of these electrons different ✓  	2	<b>ALLOW</b> other symbols for dots and crosses eg triangles  <b>IGNORE</b> charges <b>IGNORE</b> 'dative' arrow within the lone pair of the N atom
	(e) (i)	(Thermal) decomposition ✓	1	

Question		Answer	Marks	Guidance
4	(e) (ii)	<p><b>FIRST CHECK THE ANSWER ON THE ANSWER LINE</b>  <b>If answer = 242 (cm<sup>3</sup>) award 3 marks</b></p> <p>(amount of KClO<sub>3</sub>) = 0.824/122.6 <b>OR</b> = 0.00672 (mol) ✓</p> <p>(amount O<sub>2</sub>) = (mol of KClO<sub>3</sub>) 0.00672 × 3/2 <b>OR</b> = 0.0101 (mol)</p> <p>(volume of O<sub>2</sub>) = 0.0101 × 24 000 = 242 (cm<sup>3</sup>) ✓</p>	3	<p><b>IGNORE</b> over rounding to two significant figures <b>once</b>  <b>DO NOT ALLOW</b> over rounding to two significant figures twice  eg  <b>ALLOW</b> the following answer for 3 marks  <b>241</b> (cm<sup>3</sup>) (0.00672 was rounded to 0.0067 <b>OR</b> 0.0101 was rounded to 0.010)</p> <p><b>ALLOW</b> the following answers for 2 marks  <b>240</b> (cm<sup>3</sup>) (0.00672 was rounded to 0.0067 <b>AND</b> 0.0101 was rounded to 0.010)  <b>252</b> (cm<sup>3</sup>) (0.00672 was rounded to 0.007)  <b>161</b> cm<sup>3</sup> (no multiplying by 3/2)</p> <p>If there is an alternative answer, check to see if there is any ECF credit possible using working below  <b>ALLOW</b> up to correctly rounded calculator value of 0.006721044046</p> <p><b>ALLOW</b> up to correctly rounded calculator value  <b>ALLOW</b> ECF for mol of KClO<sub>3</sub> × 3/2 for 2nd mark</p> <p><b>ALLOW</b> ECF for (mol of KClO<sub>3</sub>) × 3/2 × 24000</p>
<b>Total</b>			<b>16</b>	



**OCR (Oxford Cambridge and RSA Examinations)**  
1 Hills Road  
Cambridge  
CB1 2EU

**OCR Customer Contact Centre**

**Education and Learning**

Telephone: 01223 553998

Facsimile: 01223 552627

Email: [general.qualifications@ocr.org.uk](mailto:general.qualifications@ocr.org.uk)

**[www.ocr.org.uk](http://www.ocr.org.uk)**

For staff training purposes and as part of our quality assurance programme your call may be recorded or monitored

**Oxford Cambridge and RSA Examinations**  
is a Company Limited by Guarantee  
Registered in England  
Registered Office; 1 Hills Road, Cambridge, CB1 2EU  
Registered Company Number: 3484466  
OCR is an exempt Charity

**OCR (Oxford Cambridge and RSA Examinations)**  
Head office  
Telephone: 01223 552552  
Facsimile: 01223 552553

© OCR 2012

