

Please check the examination details below before entering your candidate information

Candidate surname

Other names

Centre Number

Candidate Number

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## Pearson Edexcel Level 3 GCE

Time 1 hour 45 minutes

Paper  
reference

**9CH0/01**

# Chemistry

Advanced

**PAPER 1: Advanced Inorganic and Physical Chemistry**

**You must have:**

Scientific calculator, Data Booklet, ruler

Total Marks

### Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided  
– *there may be more space than you need.*

### Information

- The total mark for this paper is 90.
- The marks for **each** question are shown in brackets  
– *use this as a guide as to how much time to spend on each question.*
- For the question marked with an **asterisk** (\*), marks will be awarded for your ability to structure your answer logically showing the points that you make are related or follow on from each other where appropriate.
- A Periodic Table is printed on the back cover of this paper.

### Advice

- Read each question carefully before you start to answer it.
- Show all your working in calculations and include units where appropriate.
- Check your answers if you have time at the end.

Turn over ►

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Q:1/1/1/1/1/1/



P 6 7 0 9 3 R A 0 1 3 2



Pearson

**Answer ALL questions.**

**Some questions must be answered with a cross .**  
**If you change your mind about an answer, put a line through the box**   
**and then mark your new answer with a cross .**

**1** This question is about atomic structure and the Periodic Table.

(a) Complete the table to show the relative charges and masses for a neutron and an electron.

(2)

Particle	Relative charge	Relative mass
proton	+1	1
neutron		
electron		

(b) Atomic emission spectroscopy provides evidence for the existence of

(1)

- A** atoms
- B** electrons
- C** isotopes
- D** quantum shells

(c) Draw the shape of a p orbital.

(1)

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(d) The melting temperatures of two elements in Period 3 are given in the table.

Element	silicon	chlorine
Melting temperature / K	1683	172

Explain, in terms of the structure and bonding of each element, the difference between these values.

(3)

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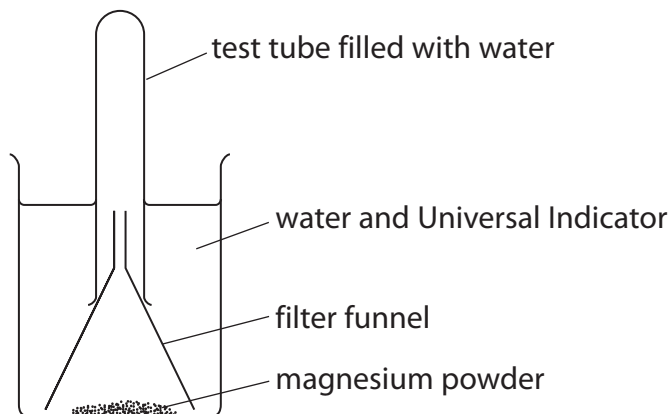
**(Total for Question 1 = 7 marks)**



2 This question is about the elements in Group 2 of the Periodic Table.

- (a) Magnesium powder is added to a beaker of water containing a few drops of Universal Indicator.

The apparatus is set up as shown and allowed to stand for a few days.



State **two** changes that will be **seen** after a few days.

(2)

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- (b) Explain how the trend in the reactivity of the Group 2 elements is determined by their electronic configurations.

(3)

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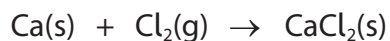
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(c) Calcium reacts with chlorine.



Explain, in terms of electrons, why this is a redox reaction.

(2)

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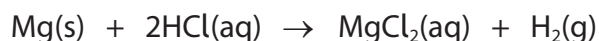
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(d) An experiment was carried out to determine the molar volume of hydrogen at room temperature.

0.035 g of magnesium was added to excess hydrochloric acid and 32 cm<sup>3</sup> of hydrogen was produced.



Calculate the molar volume of hydrogen from the results of this experiment. Include units in your answer.

(2)

**(Total for Question 2 = 9 marks)**



3 This question is about the elements in Group 7 of the Periodic Table and some of their compounds.

(a) What is the colour of iodine in the solid and gas states?

(1)

	Colour of iodine solid	Colour of iodine gas
<input type="checkbox"/> <b>A</b>	purple	brown
<input type="checkbox"/> <b>B</b>	purple	purple
<input type="checkbox"/> <b>C</b>	grey/black	brown
<input type="checkbox"/> <b>D</b>	grey/black	purple

(b) Which of these reactions occurs?

(1)

- A**  $\text{Br}_2(\text{aq}) + 2\text{NaCl}(\text{aq}) \rightarrow 2\text{NaBr}(\text{aq}) + \text{Cl}_2(\text{aq})$
- B**  $\text{Br}_2(\text{aq}) + 2\text{NaF}(\text{aq}) \rightarrow 2\text{NaBr}(\text{aq}) + \text{F}_2(\text{aq})$
- C**  $\text{Cl}_2(\text{aq}) + 2\text{NaBr}(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq}) + \text{Br}_2(\text{aq})$
- D**  $\text{Cl}_2(\text{aq}) + 2\text{NaF}(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq}) + \text{F}_2(\text{aq})$

(c) The halogens can form halide ions during reactions.

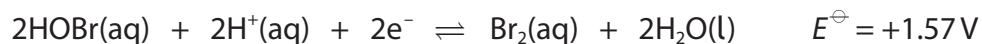
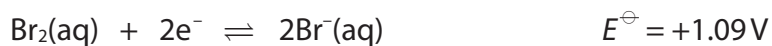
Complete the electronic configuration of the chloride ion.

(1)

$1s^2$  .....



(d) The standard electrode potentials for two half-equations involving bromine are given.



- (i) Explain why the disproportionation of bromine in water is **not** thermodynamically feasible under standard conditions. Include the overall equation for the disproportionation and its  $E_{\text{cell}}^\ominus$  value.

(3)

- (ii) Bromine disproportionates in water to a small extent at 298 K.

Give a possible reason why this reaction occurs.

(1)

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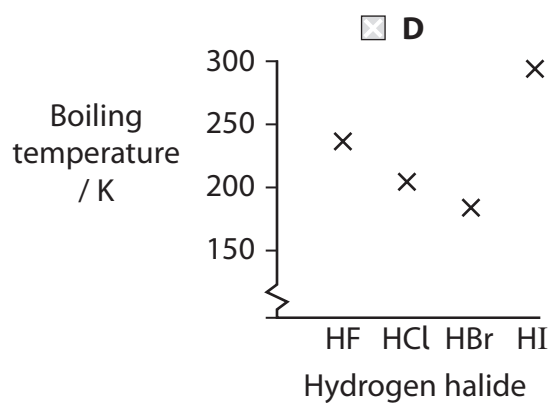
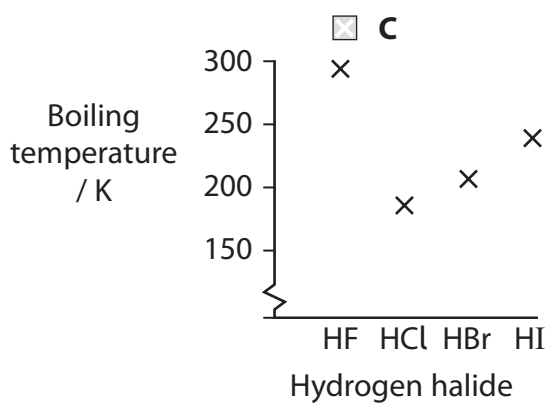
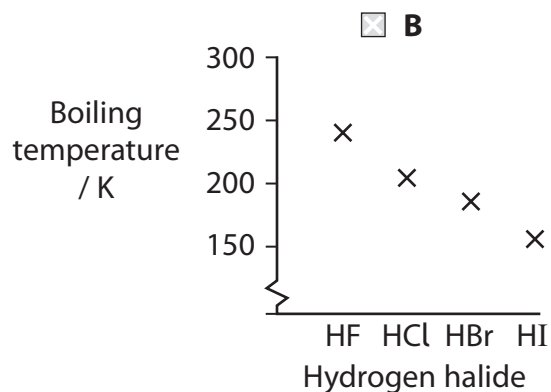
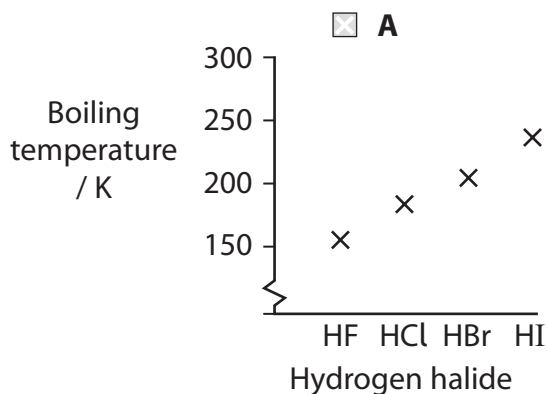


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(e) The hydrogen halides have the general formula HX, where X represents the symbol of the halogen.

(i) Which diagram shows the trend in the boiling temperatures of the hydrogen halides?

(1)



(ii) What type of reaction occurs when ammonia gas reacts with hydrogen chloride gas?

(1)

- A** acid-base
- B** displacement
- C** redox
- D** substitution

(Total for Question 3 = 9 marks)





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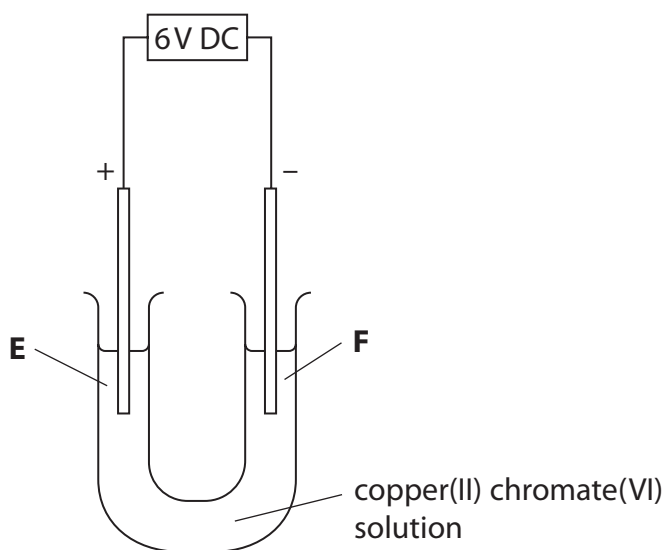
4 This question is about structure and bonding.

(a) Ionic bonding is the strong electrostatic attraction between

(1)

- A anions and cations
- B atoms and delocalised electrons
- C cations and delocalised electrons
- D two nuclei and a shared pair of electrons

(b) An aqueous solution of copper(II) chromate(VI) was electrolysed using the apparatus shown in the diagram.



Deduce the colours of the solutions in regions **E** and **F** after the electrolysis has occurred.

(2)

Colour in region **E** .....

Colour in region **F** .....



(c) Some ionic radii are given in the table.

Ion	Ionic radius / nm
Na <sup>+</sup>	0.102
Mg <sup>2+</sup>	0.072
Cl <sup>-</sup>	0.180
Br <sup>-</sup>	0.195

Deduce the **formula** of the compound, formed from the ions in the table, that has the strongest ionic bonding.

(1)

(d) The names of four substances are given.

Substance	Name
<b>P</b>	copper
<b>Q</b>	iodine
<b>R</b>	silicon(IV) oxide
<b>S</b>	sodium chloride

(i) Which of these substances exists at room temperature as a giant lattice of oppositely charged ions?

(1)

- A** Substance **P**
- B** Substance **Q**
- C** Substance **R**
- D** Substance **S**



(ii) Which of these substances has a high melting temperature, **and** conducts electricity when solid and when molten?

(1)

- A** Substance **P**
- B** Substance **Q**
- C** Substance **R**
- D** Substance **S**

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\*(e) Water has two significant anomalous properties:

- it has a higher melting temperature than hydrogen sulfide,  $\text{H}_2\text{S}$ , even though it has fewer electrons in its molecules
- the density of ice at  $0^\circ\text{C}$  is less than that of water at  $0^\circ\text{C}$ .

Explain these properties.

You should include a labelled diagram to show the intermolecular forces between two molecules of water.

(6)

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(Total for Question 4 = 12 marks)



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5 This question is about enthalpy changes and energy changes.

(a) Magnesium carbonate reacts with dilute hydrochloric acid at room temperature.



When the reaction is carried out in a sealed container with a constant volume, the heat energy change is not the same as the enthalpy change for this reaction.

Give a reason why this is so.

(1)

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(b) State what is meant by the standard enthalpy change of formation of aluminium oxide,  $\text{Al}_2\text{O}_3(\text{s})$ . Include standard conditions.

(3)

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(c) Use the data in the table to answer the questions.

Enthalpy change	Value / $\text{kJ mol}^{-1}$
Enthalpy change of hydration of $\text{K}^+$	-322
Enthalpy change of hydration of $\text{Ca}^{2+}$	-1650
Enthalpy change of solution of KCl	+17.2
Lattice energy of KCl	-711

(i) Name the two properties of ions that affect the value of their enthalpy change of hydration.

(2)

(ii) Calculate the enthalpy change of hydration for chloride ions by completing the energy cycle, including labels, and using the data in the table.

(3)



enthalpy change of hydration for  $\text{Cl}^-$  ions .....  $\text{kJ mol}^{-1}$

**(Total for Question 5 = 9 marks)**



6 This question is about acids and bases.

(a) State what is meant by a Brønsted-Lowry base.

(1)

(b) Write the ionic equation for the reaction between magnesium oxide and an acid.  
State symbols **are** required.

(2)

(c) Calculate the concentration of hydrogen ions, in  $\text{mol dm}^{-3}$ , in a solution with a pH of 9.43

(1)

(d) The pH of two salt solutions, **J** and **K**, are

solution **J**    pH = 5

solution **K**    pH = 9

The solutions are equimolar.

Which acids and bases could form the salts in solutions **J** and **K**?

(1)

	Acid and base forming the salt in solution <b>J</b>	Acid and base forming the salt in solution <b>K</b>
<input type="checkbox"/> <b>A</b>	HCl(aq) and NH <sub>3</sub> (aq)	CH <sub>3</sub> COOH(aq) and NaOH(aq)
<input type="checkbox"/> <b>B</b>	HCl(aq) and NaOH(aq)	CH <sub>3</sub> COOH(aq) and NH <sub>3</sub> (aq)
<input type="checkbox"/> <b>C</b>	CH <sub>3</sub> COOH(aq) and NaOH(aq)	HCl(aq) and NaOH(aq)
<input type="checkbox"/> <b>D</b>	CH <sub>3</sub> COOH(aq) and NH <sub>3</sub> (aq)	HCl(aq) and NH <sub>3</sub> (aq)



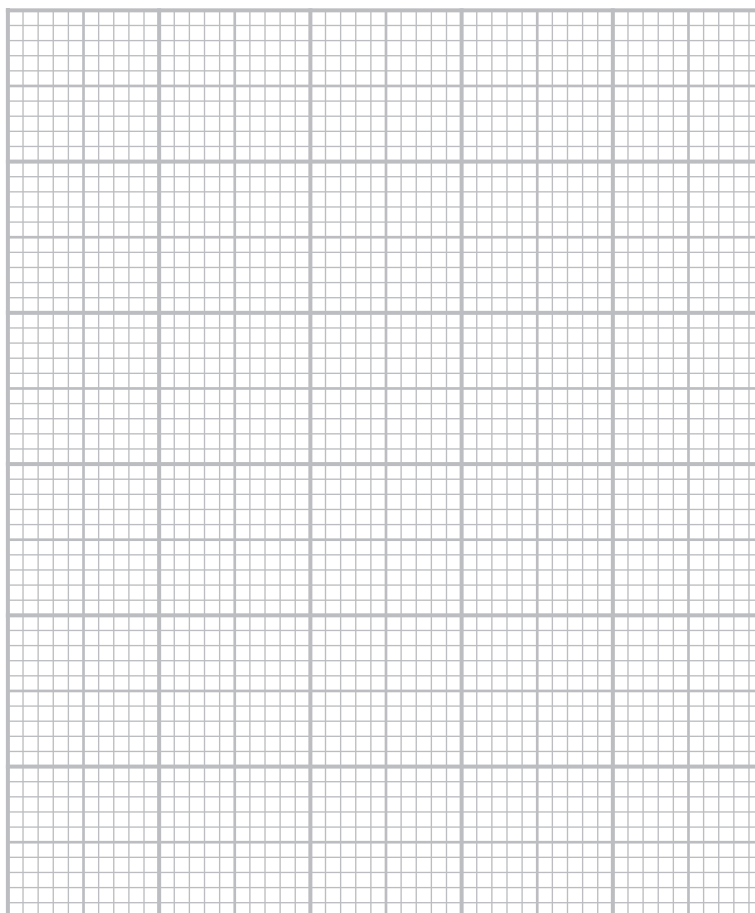


(e) The ionic product of water,  $K_w$ , varies with temperature as shown.

Temperature / °C	$K_w / \text{mol}^2 \text{dm}^{-6}$
0	$0.11 \times 10^{-14}$
10	$0.29 \times 10^{-14}$
20	$0.68 \times 10^{-14}$
30	$1.47 \times 10^{-14}$
40	$2.92 \times 10^{-14}$
50	$5.48 \times 10^{-14}$

- (i) Determine the value of  $K_w$  at 45 °C by plotting a suitable graph. You must show your working on the graph.

(3)



$K_w$  at 45 °C = .....

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(ii) The ionic product of water at 30 °C is  $1.47 \times 10^{-14} \text{ mol}^2 \text{ dm}^{-6}$ .

Calculate the pH of water at this temperature.

(3)

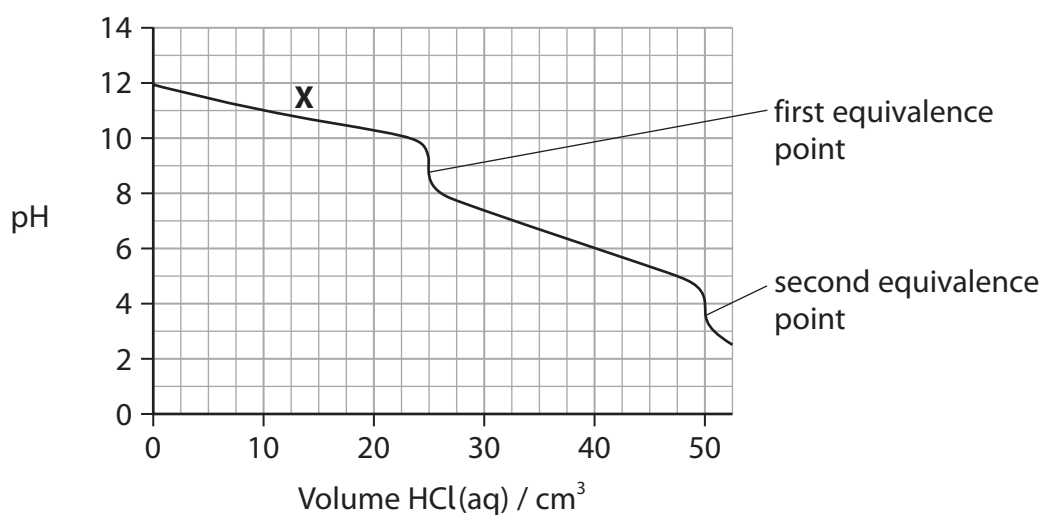
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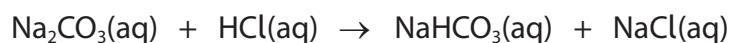
(f) Hydrochloric acid, with a concentration of  $0.100 \text{ mol dm}^{-3}$ , is added to  $25.0 \text{ cm}^3$  of  $0.100 \text{ mol dm}^{-3}$  aqueous sodium carbonate and the pH is measured.

The titration curve is shown.



The reaction takes place in two steps.

The equation for the reaction taking place in the first step is



(i) Deduce a suitable indicator to identify the first equivalence point.  
Justify your answer using values from the Data Booklet.

(2)

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(ii) Write the equation for the reaction taking place at the second equivalence point.  
State symbols are not required.

(1)

(iii) Explain how the solution at point X on the graph can act as a buffer solution.

(3)

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(Total for Question 6 = 17 marks)

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7 This question is about chromium and some of its compounds.

(a) The common oxidation numbers of chromium are +2, +3 and +6.

Give a reason, in terms of ionisation energies, why chromium can show variable oxidation numbers.

(1)

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(b) The bonding in chromate(VI) ions,  $\text{CrO}_4^{2-}$ , is similar to that in sulfate(VI) ions,  $\text{SO}_4^{2-}$ .

Draw a possible dot-and-cross diagram for a chromate(VI) ion.

(2)



- (c) A student added some pieces of zinc to an acidified solution of potassium dichromate(VI).

Some standard electrode potentials are given in the table.

Right-hand electrode system	$E^\ominus / \text{V}$
$\text{Zn}^{2+}(\text{aq}) + 2\text{e}^- \rightleftharpoons \text{Zn}(\text{s})$	-0.76
$\text{Cr}^{3+}(\text{aq}) + \text{e}^- \rightleftharpoons \text{Cr}^{2+}(\text{aq})$	-0.41
$\text{Cr}_2\text{O}_7^{2-}(\text{aq}) + 14\text{H}^+(\text{aq}) + 6\text{e}^- \rightleftharpoons 2\text{Cr}^{3+}(\text{aq}) + 7\text{H}_2\text{O}(\text{l})$	+1.33

- (i) Write the overall equation for the reduction of dichromate(VI) ions to chromium(III) ions by zinc in acid conditions.  
State symbols are not required.

(2)

- (ii) Calculate  $E_{\text{cell}}^\ominus$  for the reaction in (c)(i).

(1)

- (iii) Predict whether or not a further reduction of chromium(III) ions to chromium(II) ions will occur. Justify your answer.

(1)

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(iv) Aqueous solutions containing chromium(III) ions and chromium(II) ions have different colours.

Explain why these solutions **differ** in colour.

An explanation of the origin of the colours is not required.

(2)

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(d) An iron nail was analysed using the following outline procedure.

- An iron nail was placed in a beaker and excess dilute sulfuric acid was added.
- After all the iron had reacted to form iron(II) ions, the solution was made up to  $1.00 \text{ dm}^3$  in a volumetric flask.
- $25.0 \text{ cm}^3$  portions of the solution were acidified and titrated with potassium dichromate(VI) solution of concentration  $0.0167 \text{ mol dm}^{-3}$ .

### Results

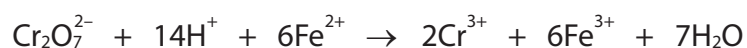
mass of iron nail = 3.54 g

mean titre =  $15.50 \text{ cm}^3$

The table shows the percentage by mass of iron in four different brands of nail.

Brand of nail	Percentage by mass of iron
<b>A</b>	92
<b>B</b>	94
<b>C</b>	96
<b>D</b>	98

Potassium dichromate(VI) in acid solution oxidises iron(II) ions as shown in the equation



Determine, using the experimental data, the brand of nail that was analysed.

(5)





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(Total for Question 7 = 14 marks)



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8 This question is about electrode potentials, cells and equilibrium constants.

(a) Chlorine gas can be prepared by the oxidation of chloride ions with manganate(VII) ions in acid solution.



During this reaction, each manganate(VII) ion accepts five electrons.

Calculate the equilibrium constant,  $K$ , for this reaction at 298 K using the expression

$$\ln K = \frac{nE_{\text{cell}}^\ominus F}{RT}$$

where  $n$  is the number of electrons transferred in the overall equation,  $F$  is the Faraday constant ( $96\,500 \text{ C mol}^{-1}$ ) and  $R$  is the gas constant ( $8.31 \text{ J mol}^{-1} \text{ K}^{-1}$ ).

Units of  $K$  are not required.

(2)



(b) A fuel cell produces a voltage from the reaction between a fuel and oxygen.

The reaction occurring at one electrode in a methanol fuel cell is



Which reaction occurs at the other electrode?

(1)

- A**  $4\text{H}^+(\text{aq}) + \text{O}_2(\text{g}) + 4\text{e}^- \rightarrow 2\text{H}_2\text{O}(\text{l})$
- B**  $2\text{H}_2(\text{g}) + 2\text{O}_2(\text{g}) + 4\text{e}^- \rightarrow 4\text{OH}^-(\text{aq})$
- C**  $4\text{OH}^-(\text{aq}) \rightarrow 2\text{H}_2(\text{g}) + 2\text{O}_2(\text{g}) + 4\text{e}^-$
- D**  $2\text{H}_2\text{O}(\text{l}) \rightarrow 4\text{H}^+(\text{aq}) + \text{O}_2(\text{g}) + 4\text{e}^-$

(c) Lead-acid batteries are used as storage cells in some cars.

The electrolyte is sulfuric acid, one electrode is lead and the other is lead(IV) oxide,  $\text{PbO}_2$ .

As the cell discharges, the lead and the lead(IV) oxide are both converted to solid lead(II) sulfate,  $\text{PbSO}_4$ , and the concentration of the sulfuric acid decreases.

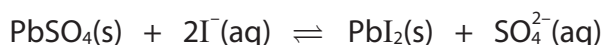
Deduce, using the information given, the two half-equations occurring in the lead-acid battery.

State symbols **are** required.

(3)



- (d) When solid lead(II) sulfate is added to aqueous sodium iodide, an equilibrium is established.



The expression for the equilibrium constant,  $K_c$ , for this reaction is

$$K_c = \frac{[\text{SO}_4^{2-}(\text{aq})]}{[\text{I}^-(\text{aq})]^2}$$

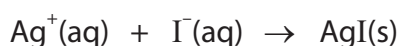
In an experiment,  $K_c$  may be determined by adding excess lead(II) sulfate to  $25.0 \text{ cm}^3$  of  $0.100 \text{ mol dm}^{-3}$  sodium iodide.

The volume remains constant at  $25.0 \text{ cm}^3$ .

The mixture is left to reach equilibrium at room temperature.

Ice-cold water is added to freeze the position of equilibrium and the mixture is then titrated with standard silver nitrate solution.

The whole mixture requires  $12.20 \text{ cm}^3$  of  $0.0500 \text{ mol dm}^{-3}$  silver nitrate solution to react with the aqueous iodide ions at equilibrium.



Calculate the equilibrium concentrations of the sulfate ions and the iodide ions, and hence the value of  $K_c$  at room temperature.

Give your answer to an appropriate number of significant figures and include units for  $K_c$ , if any.

(7)



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**(Total for Question 8 = 13 marks)**

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**TOTAL FOR PAPER = 90 MARKS**



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# The Periodic Table of Elements

1 2 3 4 5 6 7 0 (8) (18)

1.0	<b>H</b>	hydrogen	1
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## Key

relative atomic mass
<b>atomic symbol</b>
name
atomic (proton) number

(1) (2)

6.9	<b>Li</b>	lithium	3	9.0	<b>Be</b>	beryllium	4
23.0	<b>Na</b>	sodium	11	24.3	<b>Mg</b>	magnesium	12

(13)

10.8	<b>B</b>	boron	5	12.0	<b>C</b>	carbon	6	14.0	<b>N</b>	nitrogen	7	16.0	<b>O</b>	oxygen	8	19.0	<b>F</b>	fluorine	9	20.2	<b>Ne</b>	neon	10
27.0	<b>Al</b>	aluminium	13	28.1	<b>Si</b>	silicon	14	31.0	<b>P</b>	phosphorus	15	32.1	<b>S</b>	sulfur	16	35.5	<b>Cl</b>	chlorine	17	39.9	<b>Ar</b>	argon	18

(14)

69.7	<b>Ga</b>	gallium	31	72.6	<b>Ge</b>	germanium	32	74.9	<b>As</b>	arsenic	33	79.0	<b>Se</b>	selenium	34	79.9	<b>Br</b>	bromine	35	83.8	<b>Kr</b>	krypton	36
114.8	<b>In</b>	indium	49	118.7	<b>Sn</b>	tin	50	121.8	<b>Sb</b>	antimony	51	127.6	<b>Te</b>	tellurium	52	126.9	<b>I</b>	iodine	53	131.3	<b>Xe</b>	xenon	54

(15)

204.4	<b>Tl</b>	thallium	81	207.2	<b>Pb</b>	lead	82	209.0	<b>Bi</b>	bismuth	83	209.0	<b>Po</b>	polonium	84	210	<b>At</b>	astatine	85	222	<b>Rn</b>	radon	86
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(16)

112.4	<b>Cd</b>	cadmium	48	107.9	<b>Cu</b>	copper	29	63.5	<b>Zn</b>	zinc	30	65.4	<b>Zn</b>	zinc	30	107.9	<b>Ag</b>	silver	47	197.0	<b>Au</b>	gold	79	197.0	<b>Hg</b>	mercury	80
106.4	<b>Pd</b>	palladium	46	106.4	<b>Rh</b>	rhodium	45	102.9	<b>Rh</b>	rhodium	45	106.4	<b>Pd</b>	palladium	46	106.4	<b>Pt</b>	platinum	78	195.1	<b>Pt</b>	platinum	78	197.0	<b>Au</b>	gold	79

(17)

157	<b>Tb</b>	terbium	65	159	<b>Tb</b>	terbium	65	163	<b>Dy</b>	dysprosium	66	165	<b>Ho</b>	holmium	67	167	<b>Er</b>	erbium	68	169	<b>Tm</b>	thulium	69	173	<b>Yb</b>	ytterbium	70	175	<b>Lu</b>	lutetium	71
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(18)

232	<b>Th</b>	thorium	90	238	<b>U</b>	uranium	92	238	<b>U</b>	uranium	92	238	<b>U</b>	uranium	92	238	<b>U</b>	uranium	92	238	<b>U</b>	uranium	92	238	<b>U</b>	uranium	92	238	<b>U</b>	uranium	92
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Elements with atomic numbers 112-116 have been reported but not fully authenticated

\* Lanthanide series

\* Actinide series

140	<b>Ce</b>	cerium	58	141	<b>Pr</b>	praseodymium	59	144	<b>Nd</b>	neodymium	60	150	<b>Sm</b>	samarium	62	157	<b>Gd</b>	gadolinium	64	163	<b>Dy</b>	dysprosium	66	165	<b>Ho</b>	holmium	67	167	<b>Er</b>	erbium	68	169	<b>Tm</b>	thulium	69	173	<b>Yb</b>	ytterbium	70	175	<b>Lu</b>	lutetium	71				
232	<b>Th</b>	thorium	90	231	<b>Pa</b>	protactinium	91	238	<b>U</b>	uranium	92	237	<b>Np</b>	neptunium	93	242	<b>Pu</b>	plutonium	94	243	<b>Am</b>	americium	95	245	<b>Cf</b>	californium	98	251	<b>Es</b>	einsteinium	99	253	<b>Fm</b>	fermium	100	256	<b>Md</b>	mendelevium	101	254	<b>No</b>	nobelium	102	257	<b>Lr</b>	lawrencium	103

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