



Mark Scheme (Results)

January 2021

Pearson Edexcel International GCSE
In Chemistry (4CH1) Paper 2CR

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General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.

Question number	Answer	Notes	Marks														
1 (a) (i)	(from) solid to liquid		1														
	(ii) iodine (s) → iodine (g)		1														
(b)	(diagram showing) particles well spread out with none touching		1														
(c)	<table border="1" data-bbox="363 808 954 1361"> <thead> <tr> <th data-bbox="363 808 858 869">Statement</th> <th data-bbox="858 808 954 869">Tick</th> </tr> </thead> <tbody> <tr> <td data-bbox="363 869 858 929">the particles only vibrate</td> <td data-bbox="858 869 954 929"></td> </tr> <tr> <td data-bbox="363 929 858 990">the particles do not move</td> <td data-bbox="858 929 954 990"></td> </tr> <tr> <td data-bbox="363 990 858 1093">the particles have no gaps between them</td> <td data-bbox="858 990 954 1093"></td> </tr> <tr> <td data-bbox="363 1093 858 1153">the particles move randomly</td> <td data-bbox="858 1093 954 1153">✓</td> </tr> <tr> <td data-bbox="363 1153 858 1256">the particles have more energy than in ice</td> <td data-bbox="858 1153 954 1256">✓</td> </tr> <tr> <td data-bbox="363 1256 858 1361">the particles have a regular arrangement</td> <td data-bbox="858 1256 954 1361"></td> </tr> </tbody> </table> <p data-bbox="343 1400 561 1467">tick in 4th box (1) tick in 5th box (1)</p> <p data-bbox="343 1507 912 1574">if more than two ticks deduct 1 mark for each incorrect tick</p>	Statement	Tick	the particles only vibrate		the particles do not move		the particles have no gaps between them		the particles move randomly	✓	the particles have more energy than in ice	✓	the particles have a regular arrangement			2
Statement	Tick																
the particles only vibrate																	
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the particles have more energy than in ice	✓																
the particles have a regular arrangement																	

(Total for Question 1 = 5)

Question number	Answer	Notes	Marks
2 (a) (i)	(pale/light) yellow		1
	<p data-bbox="343 414 941 481">(ii) The correct answer is B 1 as bromine is the only liquid at room temperature (20 °C)</p> <p data-bbox="343 526 925 593">A 0 is incorrect because bromine is a liquid at room temperature (20 °C)</p> <p data-bbox="343 638 965 705">C 2 is incorrect because bromine is the only liquid at room temperature (20 °C)</p> <p data-bbox="343 750 973 817">D 3 is incorrect because bromine is the only liquid at room temperature (20 °C)</p>		1
	(iii) At ₂		1
(b) (i)	<p data-bbox="343 1030 622 1064">explanation including</p> <p data-bbox="343 1108 758 1142">M1 oxidising agent is chlorine/Cl₂</p> <p data-bbox="343 1176 893 1243">M2 because chlorine/Cl₂ gains electron(s)/is reduced</p>	<p data-bbox="1002 1176 1284 1310">ACCEPT because bromide ions/Br⁻ lose electrons/are oxidised REJECT bromine ions</p> <p data-bbox="1002 1355 1276 1422">M2 DEP M1 correct or missing</p>	2

Question number	Answer	Notes	Marks
2 (b) (ii)	<p>explanation containing three of following points</p> <p>M1 bromine and chlorine react by gaining electron/forming 1- or negative ion</p> <p>M2 bromine atom larger (than chlorine atom)</p> <p>M3 bromine (atom) has smaller/weaker attraction (from nucleus) for (outer shell) electrons (than chlorine) OWTTE</p> <p>M4 so (bromine has) less tendency to gain electron/form negative ion (so less reactive than chlorine) OWTTE</p>	<p>ALLOW bromine has larger atomic radius</p> <p>ALLOW bromine outer (electron) shell further from nucleus</p> <p>ALLOW bromine atom has more (electron) shells (than chlorine)</p> <p>ALLOW reverse argument for chlorine in M2 M3 M4</p>	3
(c)	<p>The correct answer is D K^+ and Cl^- because both have electronic configuration of 2.8.8</p> <p>A is not correct because Li^+ does not have electronic configuration of 2.8.8</p> <p>B is not correct because F^- does not have electronic configuration of 2.8.8</p> <p>C is not correct because neither Li^+ nor F^- have electronic configuration of 2.8.8</p>		1

(Total for Question 2 = 9)

Question number	Answer	Notes	Marks
3 (a)	Explanation including the following points: (metals) M1 contain delocalised electrons M2 (which) move/flow/are mobile/are free to move (through the metal structure) and two from (covalent compounds) M3 contain neutral molecules / molecules with no overall charge M4 electrons held (tightly) in covalent bonds (so) M5 no electrons free to move (so do not conduct)	IGNORE free electrons/sea of electrons M2 dependent on mention of electrons If any mention of ions/atoms moving cannot score M1 M2 ALLOW do not contain ions ALLOW there are no delocalised electrons If state ions present cannot score M3 M4 M5	4
(b)	ion(s)	ALLOW hydrogen ion/H ⁺ ALLOW chloride ion/Cl ⁻	1
(c) (i)	all points plotted correctly (within half a small square)		1
(ii)	point at (0.4, 25) circled		1
(iii)	straight line of best fit through origin drawn with ruler		1
(iv)	explanation linking M1 the volume/reading is less than expected		2
	M2 because the current was less than 0.4A / some gas escaped/ there was a leak	ALLOW reading taken before 5 minutes	
(v)	the greater the current the greater the volume (of gas)	ACCEPT directly proportional ACCEPT positive correlation	1

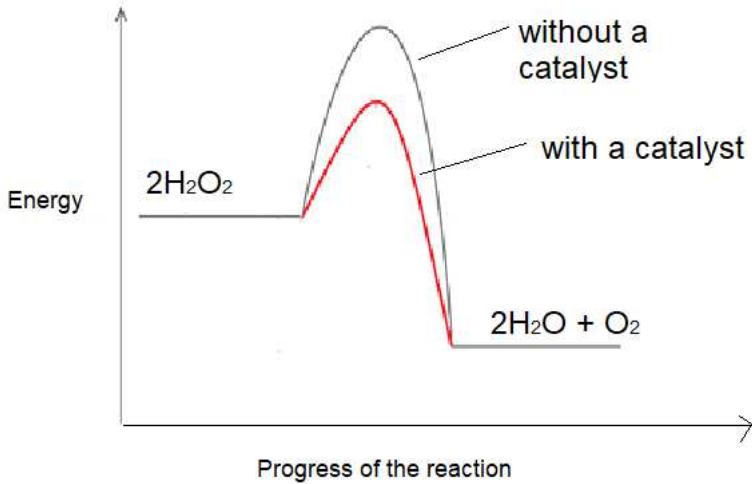
Question number	Answer	Notes	Marks
3 (d) (i)	(transfer of two/same number of electrons produces) one mole of chlorine/Cl ₂ and one mole of hydrogen/H ₂	ALLOW same number of moles of chlorine/Cl ₂ and hydrogen/H ₂ ALLOW molecules for moles	1
(ii)	chlorine dissolves/soluble in acid/solution	ALLOW dissolves/soluble in water	1

(Total for Question 3 = 13)

Question number	Answer	Notes	Marks												
4 (a)	<table border="1"> <thead> <tr> <th>Name</th> <th>Structural formula</th> <th>Relative formula mass</th> </tr> </thead> <tbody> <tr> <td>methanol</td> <td>CH₃OH</td> <td>32</td> </tr> <tr> <td>ethanol</td> <td>C₂H₅OH</td> <td>46 (1)</td> </tr> <tr> <td>butanol (1)</td> <td>C₄H₉OH</td> <td>74</td> </tr> </tbody> </table>	Name	Structural formula	Relative formula mass	methanol	CH ₃ OH	32	ethanol	C ₂ H ₅ OH	46 (1)	butanol (1)	C ₄ H ₉ OH	74	ALLOW correct names of isomers	2
Name	Structural formula	Relative formula mass													
methanol	CH ₃ OH	32													
ethanol	C ₂ H ₅ OH	46 (1)													
butanol (1)	C ₄ H ₉ OH	74													
(b) (i)	sulfuric acid	If oxidation number given must be correct ALLOW H ₂ SO ₄ IGNORE reference to concentration	1												
(ii)	from orange to green		1												
(c) (i)	ethyl ethanoate	ALLOW as one word	1												
(ii)	$\text{CH}_3\text{OH} + \text{CH}_3\text{COOH} \rightarrow \text{CH}_3\text{COOCH}_3 + \text{H}_2\text{O}$ <p>M1 CH₃COOH</p> <p>M2 CH₃COOCH₃</p>	ALLOW displayed formula	2												

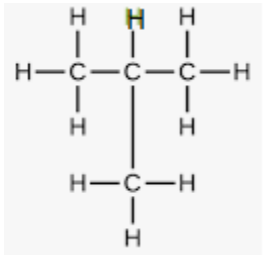
(Total for Question 4 = 7)

Question number	Answer	Notes	Marks
5 (a)	a catalyst is chemically unchanged at the end of the reaction	ALLOW (provides alternative route for reaction of) lower activation energy ALLOW not used up in reaction	1
5 (b)	<p>description including six of the following points:</p> <p>M1 do experiment using hydrogen peroxide solution only/without X/Y/Z</p> <p>M2 use known volume of hydrogen peroxide solution OWTTE</p> <p>M3 (and) measure time for certain volume of oxygen gas to be collected OR measure volume of gas collected in a certain time period OWTTE</p> <p>M4 repeat using same volume of hydrogen peroxide solution</p> <p>M5 with known mass/amount of solid X (then Y, then Z)</p> <p>M6 measure time for same volume of oxygen gas to be collected OR measure volume of gas collected in same time period (with solid/X/Y/Z present)</p> <p>M7 after reaction (remove solid/X/Y/Z by filtration and dry) find mass of solid/X/Y/Z /check if mass unchanged</p> <p>M8 reference to reduced time (for certain volume of oxygen gas to be collected) OR increased volume of gas (collected in a certain time period) means X/Y/Z (possible) catalyst (1)</p>	<p>ALLOW measure time until no more oxygen produced</p> <p>ALLOW amount</p> <p>ALLOW reference to increased rate</p>	6

Question number	Answer	Notes	Marks
5 (c)	<p>M1 labelled profile curve drawn for reaction without catalyst</p> <p>M2 labelled profile curve drawn with lower activation energy for reaction with catalyst</p> 	must start from reactants energy level and end at products energy level	2
(d) (i)	energy needed = $[(4 \times 463) + 2(146)] = 2144$	Ignore sign	1
(ii)	energy released = $[(4 \times 463)] = 1852$	ignore sign	1
(iii)	<p>M1 for showing equation linking between (i), (ii), O=O bond energy and ΔH</p> <p>M2 correct calculation</p> <p>Example:</p> <p>M1 $[1852 + (\text{O}=\text{O})] - 2144 = 204$</p> <p>OR $(\text{O}=\text{O}) = 2144 - 1852 + 204$</p> <p>M2 $(\text{O}=\text{O}) = 496$</p>	<p>Mark CSQ on (i) and (ii)</p> <p>-496 scores 1 mark</p>	2

(Total for Question 5 = 13)

Question number	Answer	Notes	Marks												
6 (a) (i)	pipette	ALLOW graduated pipette	1												
	(ii) The correct answer is D yellow because methyl orange is yellow in alkaline solution A is incorrect because methyl orange is not blue in alkaline solution B is incorrect because methyl orange is not orange in alkaline solution C is incorrect because methyl orange is not red in alkaline solution		1												
(b) (i)	ticks under titration numbers 1, 2 and 4 <table border="1" data-bbox="363 846 975 1032"> <thead> <tr> <th>1</th> <th>2</th> <th>3</th> <th>4</th> </tr> </thead> <tbody> <tr> <td>20.65</td> <td>20.60</td> <td>20.90</td> <td>20.55</td> </tr> <tr> <td>✓</td> <td>✓</td> <td></td> <td>✓</td> </tr> </tbody> </table>	1	2	3	4	20.65	20.60	20.90	20.55	✓	✓		✓		1
1	2	3	4												
20.65	20.60	20.90	20.55												
✓	✓		✓												
	(ii) M1 $\frac{20.65 + 20.60 + 20.55}{3} =$ M2 20.60	M1 CSQ on results ticked M2 CSQ on correct calculation from M1 Answer to M2 must be correct to 2dp 20.60 without working scores 2 20.6 with or without working scores 1 If no ticks then only use of 2 or 3 concordant titres can score M1 and M2 in (ii) If only one tick then M2 can be scored for averaging two or more titre values correctly	2												

Question number	Answer	Notes	Marks
7 (a)	<p>M1 use of amount in moles = $\frac{\text{volume of gas}}{\text{molar volume}}$</p> <p>M2 correct evaluation</p> <p>Example calculation</p> <p>M1 amount in moles = $\frac{600}{24000}$</p> <p>M2 = 0.025 (mol)</p>	<p>0.025 with no working scores 2 REJECT 0.03 for M2</p>	2
(b)	<p>M1 statement/use of amount in moles = $\frac{\text{mass}}{Mr}$</p> <p>M2 substitution and correct evaluation:</p> $Mr = \frac{1.45}{0.025} = 58$		2
(c)	<p>molecular formula = $C_4H_{10} = (4 \times 12) + (10 \times 1) = 58$</p> <p>OR</p> <p>alkane general formula = C_nH_{2n+2}</p> <p>so</p> <p>M1 $(n \times 12) + (2n+2) \times 1 = 58$ so $14n = 56$</p> <p>M2 $n = 4$ so C_4H_{10}</p>		1
(d)	 <p>The diagram shows the structural formula of butane, a four-carbon alkane. The carbon atoms are arranged in a horizontal chain, with the second carbon atom from the left highlighted in yellow. Each carbon atom is bonded to hydrogen atoms to satisfy its four bonds: the first and last carbons are bonded to three hydrogens each, and the two middle carbons are bonded to two hydrogens each.</p>		1

(Total for Question 7 = 6)

Question number	Answer	Notes	Marks
8 (a) (i)	$\text{NH}_4\text{Cl} + \text{NaOH} \rightarrow \text{NaCl} + \text{NH}_3 + \text{H}_2\text{O}$		1
(ii)	M1 moist/damp red litmus paper M2 turns blue	ALLOW moist/damp universal indicator paper	2
(b)	an explanation linking following two points: M1 forward and backward/reverse reactions are taking place at same rate M2 the concentrations of reactants and products remain constant	ACCEPT reaction is taking place in both directions at same rate REJECT concentrations of the reactants and products are equal/the same	2
(c)	an explanation linking following two points: M1 as temperature decreases yield of ammonia (formed in forward reaction) increases M2 so forward reaction is exothermic	ALLOW as temperature decreases equilibrium position shifts in forward direction/(from left) to right (producing more ammonia) IGNORE references to pressure ALLOW reverse arguments IGNORE references to Le Chatelier's Principle M2 DEP M1	2

(Total for Question 8 = 7)

Total for Paper = 70 marks

